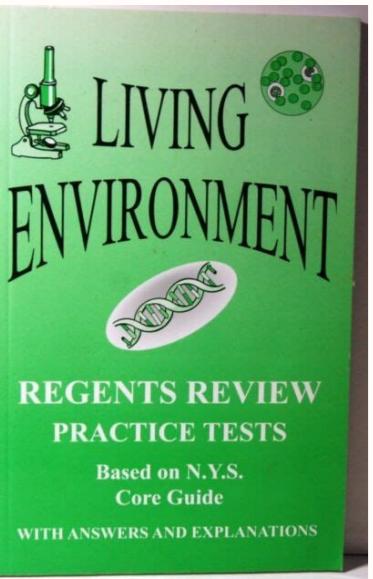


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Living environment regents review questions and answers



Part A

Answer all questions in this part.

Directions (1–30): For each statement or question, record on your separate answer sheet the number of the word or expression that, of those given, best completes the statement or answers the question. Some questions

- may require the use of the 2011 Edition Reference Tables for Physics, Chemistry, and Mathematics.

1 Compared to the charge of a proton, the charge of an electron has
(1) a greater magnitude and the same sign
(2) a greater magnitude and the opposite sign
(3) the same magnitude and the same sign
(4) the same magnitude and the opposite sign

2 Which atom has the largest atomic radius?
(1) potassium 200 (3) francium 242
(2) rubidium 215 (4) cesium 238

3 In the wave-mechanical model of the atom, an orbital is defined as
(1) a region of the most probable proton location
(2) a region of the most probable electron location
(3) a circular path traveled by a proton around the nucleus
(4) a circular path traveled by an electron around the nucleus

4 When an excited electron in an atom moves to the ground state, the electron
(1) absorbs energy as it moves to a higher energy state
(2) absorbs energy as it moves to a lower energy state
(3) emits energy as it moves to a higher energy state
(4) emits energy as it moves to a lower energy state

5 Which polyatomic ion is found in the compound represented by the formula NaHCO_3 ?
(1) acetate (3) hydrogen sulfate
(2) hydrogen carbonate (4) oxalate

6 The atomic mass of magnesium is the weighted average of the atomic masses of
(1) all of the artificially produced isotopes of Mg
(2) all of the naturally occurring isotopes of Mg
(3) the two most abundant artificially produced isotopes of Mg
(4) the two most abundant naturally occurring isotopes of Mg

#24 7 Which element has atoms that can form halide ions?
(1) iodine (3) strontium
(2) silver (4) xenon

#1 8 Two forms of solid carbon, diamond and graphite, differ in their physical properties due to the differences in their
(1) atomic numbers
(2) crystal structures
(3) isotopic abundances
(4) percent compositions

#4 9 Which quantity can be calculated for a solid compound, given only the formula of the compound and the Periodic Table of the Elements?
(1) the density of the compound
(2) the heat of fusion of the compound
(3) the melting point of each element in the compound
(4) the percent composition by mass of each element in the compound

10 Which terms identify types of chemical reactions?
(1) decomposition and sublimation
(2) decomposition and synthesis
(3) deposition and sublimation
(4) deposition and synthesis

P.S./Chem.-June '34

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