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The text still contains unprecedented macroscopic-to-microscopic molecular illustrations, consistent step-by-step worked exercises in every chapter, and an extensive range of end-of-chapter problems, which provide engaging applications covering a wide variety of interests, including engineering, medicine, materials, and environmental studies. Changes have been made to the text and applications throughout to make them more succinct, to the artwork to make it more teachable and modern, and to the design to make it more simplistic and open. Free download Chemistry The Molecular Nature of Matter and Change with Advanced Topics (8th Edition) written by Martin S. Silberberg and Patricia G. Amateis in pdf. The eighth edition of Chemistry: The Molecular Nature of Matter and Change maintains its standard-setting position among general chemistry textbooks by evolving further to meet the needs of professor and student. The text still contains the most accurate molecular illustrations, consistent step-by-step worked problems, and an extensive collection of end-of-chapter problems. And changes throughout this edition make the text more readable and succinct, the artwork more teachable and modern, and the design more focused and inviting. The three hallmarks that have made this text a market leader are now demonstrated in its pages more clearly than ever. A favorite feature, the section summaries that conclude every section restate the major ideas concisely and immediately (rather than postponing such review until the end of the chapter). A rich catalog of study aids ends each chapter to help students review the content: · Learning Objectives, with section and/or sample problem numbers, focus on the concepts to understand and the skills to master. · Key Terms, boldfaced and defined within the chapter, are listed here by section (with page numbers), as well as being defined in the Glossary. · Key Equations and Relationships are highlighted and numbered within the chapter and listed here with page numbers. · Brief Solutions to Follow-up Problems triple the number of worked problems by providing multi-step calculations at the end of the chapter, rather than just numerical answers at the back of the book. Contents Keys to Studying Chemistry: Definitions, Units, and Problem Solving The Components of Matter Stoichiometry of Formulas and Equations Three Major Classes of Chemical Reactions Gases and the Kinetic-Molecular Theory Thermochemistry: Energy Flow and Chemical Change Quantum Theory and Atomic Structure Electron Configuration and Chemical Periodicity Models of Chemical Bonding The Shapes of Molecules Theories of Covalent Bonding Intermolecular Forces: Liquids, Solids, and Phase Changes The Properties of Mixtures: Solutions and Colloids Periodic Patterns in the Main-Group Elements Organic Compounds and the Atomic Properties of Carbon Kinetics: Rates and Mechanisms of Chemical Reactions Equilibrium: The Extent of Chemical Reactions Acid-Base Equilibria Ionic Equilibria in Aqueous Systems Thermodynamics: Entropy, Free Energy, and Reaction Direction Electrochemistry: Chemical Change and Electrical Work The Elements in Nature and Industry Transition Elements and Their Coordination Compounds Nuclear Reactions and Their Applications Free download Chemistry The Molecular Nature of Matter and Change with Advanced Topics 8th Edition by Martin S. 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The revision for the eighth edition focused on continued optimization of the text. .To aid in this process, we were able to use data from literally thousands of student responses to questions in LearnSmart, the adaptive learning system that assesses student knowledge of course content. .The data, such as average time spent answering each question and the percentage of students who correctly answered the question on the first attempt, revealed the learning objectives that students found particularly difficult, which we addressed by revising surrounding text or adding additional learning resources such as videos and slideshows. The text still contains unprecedented macroscopic-to-microscopic molecular illustrations, consistent step-by-step worked exercises in every chapter, and an extensive range of end-of-chapter problems, which provide engaging applications covering a wide variety of interests, including engineering, medicine, materials, and environmental studies. Changes have been made to the text and applications throughout to make them more succinct, to the artwork to make it more teachable and modern, and to the design to make it more simplistic and open. .Sample questions asked in the 8th edition of Chemistry: The Molecular Nature of Matter and Change.Using 21 st -century technology, hydrogen fusion requires temperatures around 10 8 K. But, lower initial temperatures are used if the hydrogen is compressed. In the late 24 th century, the starship Leinad uses such methods to fuse hydrogen at 10 6 K. (a) What is the kinetic energy of an H atom at 1.00×10 6 K? (b) How many H atoms are heated to 1.00×10 6 K from the energy of one H and one anti-H atom annihilating each other? (c) If the heated H atoms of part (b) fuse into 4 He atoms (with the loss of two positrons per 4 He formed), how much energy (in J) is generated? (d) How much more energy is generated by the fusion in part (c) than by the hydrogen-anthhydrogen collision in part (b)? (e) Should the captain of the Leinad change the technology and produce 3 He (mass = 3.01603 amu) instead of 4 He?Consider the following balanced thermochemical equation for a reaction sometimes used for H 2 S production: ? (a) Is this an exothermic or endothermic reaction? (b) What is ? H for the reverse reaction? (c) What is ? H when 2.6 mol of S 8 reacts? (d) What is ? H when 25.0 g of S 8 reacts?The compound dimethyl phthalate, used in insect repellants, is composed of carbon, hydrogen, and oxygen and has a molar mass of 194.2 g/mol. A 0.2572-g sample of the compound was burned in a combustion apparatus and 0.583 g of CO 2 and 0.119 g of H 2 O were collected. What is the molecular formula of dimethyl phthalate?At death, a nobleman in ancient Egypt was mummified and his body contained 1.4×10 ?3 g of 40 K (t 1/2 = 1.25×10 9 yr), 1.2×10 ?8 g of 14 C (t 1/2 = 5730 yr), and 4.8×10 ?14 g of 3 H (t 1/2 = 12.26 yr). Which nuclide would give the most accurate estimate of the mummy's age? 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